

CH 105 Supplemental Instruction

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Sessions: Monday, 1:15-2:15, EB 128

Wednesday, 3:30-4:30, EB 133

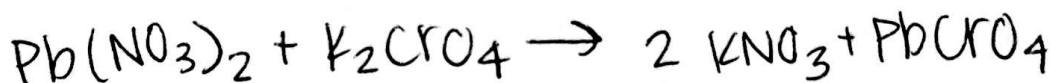
Office Hour: Thursday, 3:30-4:30, EB 242 (Academic Success Center)

1. Reaction Practice: Write a balanced equation for the following reactions:

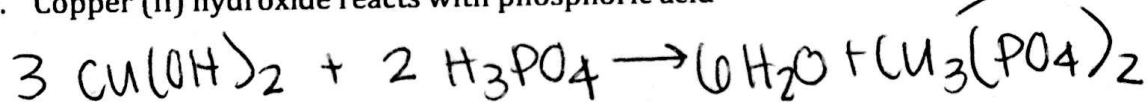
a. ~~Calcium oxide~~ reacts with hydrofluoric acid
Calcium hydroxide



b. Lead (II) nitrate reacts with potassium chromate



c. Copper (II) hydroxide reacts with phosphoric acid



d. Zinc reacts with ~~bicarbonate~~. carbonic acid



2. Which reactions above are acid-base reactions? How do you know? How do you identify a base and an acid?

A + C. They have an acid & base in the reactants. Acids have H in the beginning and bases have OH.

3. In each of the acid base reactions above, label the salt.

4. The acid and base neutralize each other and a salt is formed.

5. What are the products of complete burning or oxidation ALWAYS?? What elements must be in the ~~products~~ ^{reactants} for this to happen?

H₂O & CO₂. An organic compound must be in the reactants. ~~organic compounds have C + H. O₂ is in air.~~
organic compounds have C + H. O₂ is in air.

6. Write down the number of a mole. Memorize and love this number. You're gonna need it! ☺

$$6.02 \times 10^{23}$$