CH 105 Supplemental Instruction

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Sessions: Monday, 1:15-2:15, EB 128
Wednesday, 3:30-4:30, EB 133
Office Hour: Thursday, 3:30-4:30, EB 242 (Academic Success Center)

Session 5

1. Elements are made of one type of atom but can have more than one atom. * but remember w/ isotopes, atoms can be different
2. True or (False): there is something simpler than an element in chemical composition.
3. In the periodic table, columns are called groups and rows are called periods.
4. In which of these, columns or rows, do the elements have similar characteristics?
   groups

Make sure that you can label the main elements and transition elements on a periodic table.

5. Name the common groups of elements.
   a. 1A: Alkali metals
   b. 2A: Alkaline earth metals
   c. 3A-6A: ~~transition elements~~
   d. 7A: Halogens
   e. 8A: Noble gases

6. What is Dalton's Atomic Theory? Briefly touch on the rules.
   - tiny particles called atoms
   - atoms are indivisible; they cannot be destroyed [not entirely true]
   - all atoms of one element are identical in every respect *
   - atoms are different from atoms of any other element
   - atoms combine to form chemical compounds.

7. Who discovered the electron? Who discovered the proton? The neutron? Which one is positive and which one is negative?
   electron: Michael Faraday + Williams Crooke
   * remember J.J. Thompson
   proton: Rutherford
   neutron: Chadwick
   proton = positive
   electron = negative

8. What happens when atoms gain or lose electrons?

ions are formed

9. Describe the Plum Pudding model of the atom.

Negative e- floating in positive "pudding". Thompson is responsible for this.